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$226\ 88\ \text{Ra} \rightarrow 222\ 86\ \text{Rn} + 4\ 2\ \alpha$. Because nucleons are conserved in this and all other nuclear reactions, the sum of the mass numbers of the products, $222 + 4 = 226$, equals the mass number of the parent. Similarly, the sum of the atomic numbers of the products, $86 + 2 = 88$, equals the atomic number of the parent.

Chapter 25

804 Chapter 25 Nuclear Chemistry CHAPTER 25 What You'll Learn You will trace the history of nuclear chemistry from discovery to application. You will identify types of radioactive decay and solve decay rate problems. You will describe the reactions involved in nuclear fission and fusion. You will learn about applications of nuclear reactions and the effects of radiation exposure.

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radioactive decay, is an example of a nuclear reaction. • Nuclear reactions begin with unstable isotopes, or radioisotopes. • Atoms of these isotopes become more stable when changes occur in their nuclei. • The changes are always accompanied by the emission of large amounts of energy. 25.1 Nuclear Radiation >

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Amour in your notebook solve the following problems section 251 nuclear radiation 1 what happens to the

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Mass number and atomic number are conserved in all nuclear reactions. The mass of a 25.0 g piece of ^{238}Cm (half-life: 2.4 hr) will be reduced to 147 3.1 g after 7.2 hr. Study Guide for Content Mastery Chemistry: Matter and Change Chapter 25

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800 Chapter 25 Types of Radiation Discuss Explain that the nuclei of a radioactive element spontaneously decompose. Nuclear chemistry is the study of changes in matter that originate in atomic nuclei. Ask, What types of radiation exist, and how harmful are they? (The three most common types of radiation emitted by unstable nuclei are

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