

Hydrated Crystal Lab Answers

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HYDRATED COMPOUNDS LAB

Solved: Lab 09 - Percent Of Water In A Hydrate Pre-Lab Que ...

4 is the right substance for the moisture absorber and indicator, you will need to examine the hydrated and anhydrous forms of the compound and determine the following:

- the empirical formula of the hydrate, including its water of crystallization,
- if the compound is useful as an indicator when it changes from the hydrated to the anhydrous form, and
- the mass of water absorbed by the 25 g of anhydrous compound, which the company proposes to use.

Hydrate Lab - Google Docs

The crystals change form, and sometimes color, as the water is driven off. This suggests that water was present as part of the crystal structure. Such compounds are called hydrates. A hydrate that has lost its water is called an anhydrous salt. For a hydrate, the number of moles of water present per mole of salt is usually some simple, whole number.

EXPERIMENT 7: HYDRATES Introduction: Background

the crystal structure. Crystalline compounds that retain water during evaporation are referred to as being hydrated or are said to contain water of hydration. The ratio of moles of water to moles of compound is a small whole number. Example: $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ Mole ratio: 1:2 Name: barium chloride dihydrate THIS LAB: $\text{CuSO}_4 \cdot ?\text{H}_2\text{O}$

Hydrated Crystals Lab? | Yahoo Answers

For the mass of water in hydrated MgSO_4 , just subtract the mass of the anhydrous MgSO_4 from the mass of the hydrated MgSO_4 (the difference must be the water). For moles anhydrous MgSO_4 , divide the...

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Hydrated Crystals Teacher Information Lab Overview In this lab students will heat magnesium sulfate hydrate (Epsom salts) until the water has been removed. They can then calculate the percent water in the hydrate and use that to calculate the formula of the hydrated crystal and compare it to theoretical data.

Rancho High School

1. Compare the number of moles of anhydrous MgSO_4 , to the number of moles of water in the hydrate. Use the ratio of these two values to predict a formula for hydrated MgSO_4 . 2. The method used in this...

Formula of a Hydrate Lab

Heat hydrate crystals to the crucible. Determine the mass of ... before you leave the lab. Data Table
Mass of empty crucible and cover 39
Mass of crucible, cover, and magnesium sulfate ...
Could the solid be a hydrate? What evidence supports your answer?
T b. If, after heating, the solid has a molar mass of 208 g/mol and a formula of ...

Hydrated Crystal Lab Answers

These compounds are referred to as being “hydrated” (meaning, to contain water) and are said to have a “water of hydration”. These water molecules are scattered within the crystal structure rather loosely; therefore heating over a low heat will drive off the water molecules.

2020_LAB_3-_Hydrate_Lab (1).doc - AP Chemistry Lab#2 ...

Hydrate Lab. Kimberly Graziano & Hyunjae Kim. Purpose. Experimental Question: How can we experimentally determine the formula of an unknown hydrate, A? Testable Prediction: Our unknown hydrate may be a hydrate of copper(II) sulfate, magnesium sulfate, iron(III) chloride, or iron(III) nitrate. Observing our nitrate, it has a white crystalline structure, representing that similar to table salt.

WordPress for Faculty & Staff - at Clark University

Your Hydrated Crystal formula: $\text{NiSO}_4 \cdot 6\text{H}_2\text{O}$. Complete the following calculations. 1) Calculate the total molar mass of the formula (remember to use the number in front of the H_2O). 2) Calculate the molar mass of just the water in the formula (Use the multiple in front of the H_2O). 3) Calculate the % composition

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of water in the

LAB: Percent Composition of Hydrated Crystals

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Hydrate Anhydride + Water . $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ $\text{CaSO}_4(\text{s}) + 2\text{H}_2\text{O}(\text{g})$ Therefore, you can determine the percentage of water in a hydrate by determining the mass lost (amount of water driven off) when a known mass of hydrate is heated.
Percentage of water = $\frac{\text{Mass of water lost}}{\text{Mass of hydrate}} \times 100$

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Discussions: 3. Theoretical value of Mass % of Crystalline water in hydrous salt = $\frac{\text{Mass of water present in hydrated crystals} \times 100}{\text{Molar mass of hydrated salt}}$ For $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$; Theoretical value of Mass view the full answer

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Copper Sulfate's Water of Hydration Lab

Hydrates are solid ionic compounds (salts) that contain water molecules as part of their crystal structure. Ex: $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ (s) is a hydrate that contains one copper (II) formula unit with 5 molecules of water attached. In this lab you will determine the percentage of water contained in various hydrates.

Percent Composition of Hydrates

The formula for the hydrated crystal tells how many water molecules are attached to each formula unit of crystal. For example, hydrated calcium nitrate is written: $\text{Ca}(\text{NO}_3)_2 \cdot 4 \text{H}_2\text{O}$ and is called calcium nitrate tetrahydrate. The dot in the formula shows that for every mole of calcium nitrate in the hydrated crystal, there are 4 moles of water. The water is attached physically to the crystal and can be removed from most hydrated crystals by heating them. The crystal with the water removed ...

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